2011

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Antiaromatic Dianions: Dianions of Dixanthylidene by Reduction and Attempted Excited-State Deprotonation

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Supporting Information

ABSTRACT: Reduction of dixanthylidene with potassium or lithium resulted in formation of the antiaromatic dianion in high yield. Attempts to form the dianion by excited-state deprotonation of dixanthene with n-butyllithium/TMEDA resulted in formation of the tetraanion from deprotonation ortho to the oxygen. Orientation of the sp^3 hydrogens presumably allows preferential deprotonation of the xanthene rings.

We have been involved in the preparation of antiaromatic dications via oxidation of unsaturated precursors and ionization of diols for some time. The antiaromaticity of these species can be evaluated by the same type of criteria used in the evaluation of aromatic compounds. Those criteria are magnetic, including ^1H NMR shifts, nucleus-independent chemical shifts, and magnetic susceptibility exaltation; energetic, including aromatic stabilization energy, and structural, including the degree of bond length alternation. While these criteria have not always provided consistent answers to the relative aromaticity of compounds, the inclusion of antiaromatic species can provide greater clarity to the relationships between the criteria because the range of measurements becomes larger. Focusing only on magnetic measures, the proton chemical shifts for neutral aromatic species are usually found between 6.5 and 8.5 ppm; the addition of antiaromatic dications can give ranges beginning as high as 4.97 ppm; antiaromatic dianions give ranges which begin as high as 2.58 ppm. Magnetic susceptibility exaltation is known to be dependent on the area of the ring system; the extension of range provided by the inclusion of antiaromatic species has allowed the relationship between magnetic susceptibility exaltation and the nucleus independent chemical shift to confirm a similar dependence on the area of the ring system for the nucleus independent chemical shift.

While the bulk of our work has been on antiaromatic dications, we have begun to explore the antiaromaticity of dianions such as the dianion of tetrabenzo[5.7]fulvalene, which we prepared by reduction of the neutral precursor. The success of this mode of formation suggested that the following suite of dianions, 1^-^-^-^- and monoanion, e.g., 4, also possess planar ring systems; see the Supporting Information.

Measures of Antiaromaticity. Nucleus-Independent Chemical Shift. The nucleus-independent chemical shifts, NICS, provide a measure of aromaticity/antiaromaticity through a probe of the magnetic region in the center of a ring current. The calculated chemical shift of this probe uses the tensor perpendicular to the planar ring system and is known as the NICS(1)_zz value. Negative NICS values indicate aromaticity; positive values, antiaromaticity. The NICS(1)_zz values for dianions 1^-^-^- and monoanion 4^- are given in Table 1. Thus, the all-carbon system 1^-^-^ is the most antiaromatic, with the antiaromaticity decreasing as the heteroatom becomes more electronegative.

Magnetic Susceptibility Exaltation. A second measure of aromaticity/antiaromaticity is the exaltation of the magnetic susceptibility that is caused by the response of the ring current to the magnetic field. The magnetic susceptibility exaltation, Δ, can be determined by subtracting the sum of the magnetic susceptibilities of the bonds in a molecule or ion from the magnetic susceptibility of the delocalized system. The exaltation for 1^-^-^-^- is given in Table 1 and shows the same...
decrease in antiaromaticity shown by the \( \Sigma \text{NICS}(1)_{zz}/\text{square area} \), see the Supporting Information for details of the calculations. We have previously shown a linear relationship between \( \Sigma \text{NICS}(1)_{zz}/\text{square area} \) and \( \Lambda/\text{square area} \). That plot is given in Figure 1 for a series of aromatic and antiaromatic ions and neutral compounds. Dianions \(^1\)\(^2\)\(^3\) and monoanion \(^4\) show the same relationship as the larger group of neutral compounds and anions that was taken from the previous study.

Comparison of the Antiaromaticity of Dianion \(^3\)\(^2\) with Monoanion \(^4\). Because \(^3\)\(^2\)\(^1\) is the first antiaromatic heterocyclic dianion for which structural evidence exists, it is appropriate to compare it with the corresponding monoanion \(^4\). \(^1\)H NMR spectral data for \(^4\) has been reported\(^{26,27}\) and is included in the Supporting Information. The average \(^1\)H chemical shift for \(^4\) from deprotonation with KH in DMSO-\(d_6\) is 5.428 ppm; that of \(K_2\)\(^3\) is 5.242 ppm, vide infra. While it is difficult to compare chemical shifts in different solvents, this data suggests slightly greater antiaromaticity for the dianion over the corresponding monoanion. This is consistent with the increased antiaromaticity of the dication of tetrabenzo[5,5]fulvalene,\(^6\) in comparison with the corresponding fluorenyl monocation.\(^{28}\) The calculated NICS values of dianion \(^3\)\(^2\)\(^1\) were more positive, supporting its greater antiaromaticity than those of the corresponding monoanion \(^4\), when evaluated per ring system. It appears that this relationship is not consistent with the magnetic susceptibility data of \(^4\) compared to a single ring system of \(^3\)\(^2\). However, Figure 1 suggests that this inconsistency does not affect the relationship between the \( \Sigma \text{NICS}/\text{square area} \) and \( \Lambda/\text{square area} \). A probable explanation lies in the relatively small degree of antiaromaticity in both \(^3\)\(^2\) and \(^4\) in comparison with other antiaromatic dications and dianions studied.

Preparation of the Dianion of Dixanthylidene by Reduction. Reduction of dixanthylidene, \(^3\), with either lithium or potassium resulted in the formation of \(^3\)\(^2\); see Scheme 1.

The spectrum of \(^3\)\(^2\) in THF-\(d_8\) from reduction with potassium is shown in Figure 2a and that with lithium in Figure 2b. For the complete \(^1\)H NMR spectra from each reduction and the COSY spectrum from reduction with potassium, see the Supporting Information. While the chemical shifts of both \(K_2\)\(^3\) and \(Li_2\)\(^3\) show the upfield shift consistent with an antiaromatic species, anionic species also show upfield \(^1\)H NMR chemical shifts. To understand the contribution of antiaromaticity to the \(^1\)H NMR chemical shifts, we calculated the chemical shifts using the GIAO method that was also used for the calculation of the NICS values. The chemical shifts were also calculated with solvent using the polarization continuum, PCM, method, and with the

Table 1. Nucleus-Independent Chemical Shifts\(^a\), \( \Sigma \text{NICS}(1)_{zz}\), and Magnetic Susceptibility Exaltation\(^b\), \( \Lambda \), for \(^1\)\(^2\)\(^3\) and \(^4\).

<table>
<thead>
<tr>
<th></th>
<th>central ring</th>
<th>outer rings</th>
<th>( \Sigma \text{NICS}(1)_{zz} )</th>
<th>( \Lambda )</th>
</tr>
</thead>
<tbody>
<tr>
<td>(^1)(^2)(^3)</td>
<td>104.67</td>
<td>38.07</td>
<td>361.60</td>
<td>236.14</td>
</tr>
<tr>
<td>(^2)(^3)</td>
<td>23.81</td>
<td>11.28</td>
<td>112.76</td>
<td>75.43</td>
</tr>
<tr>
<td>(^3)(^2)(^1)</td>
<td>32.77</td>
<td>1.04</td>
<td>69.68</td>
<td>33.02</td>
</tr>
<tr>
<td>(^4)</td>
<td>9.87</td>
<td>0.65</td>
<td>11.17</td>
<td>24.37</td>
</tr>
</tbody>
</table>

\(^a\) Calculated with the GIAO method and basis set B3LYP/6-311+g(dp) on geometries optimized at the B3LYP/6-31 g(d) level. \(^b\) Calculated with the CSGT method and basis set B3LYP/6-311+g(dp) on geometries optimized at the B3LYP/6-31 g(d) level. \(^c\) Summation for entire system.
Table 2. Experimental and Calculated Shifts for K$_3$3 and for Li$_3$3

<table>
<thead>
<tr>
<th></th>
<th>K$_3$3</th>
<th>Li$_3$3</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td>expt</td>
<td>calc$^a$</td>
</tr>
<tr>
<td>1</td>
<td>4.664</td>
<td>4.425</td>
</tr>
<tr>
<td>2</td>
<td>5.921</td>
<td>6.136</td>
</tr>
<tr>
<td>3</td>
<td>5.144</td>
<td>5.434</td>
</tr>
<tr>
<td>4</td>
<td>5.240</td>
<td>5.462</td>
</tr>
</tbody>
</table>

$^a$ Shifts calculated with the GIAO and PCM methods with basis set B3LYP/6-311+g(d,p) on geometries optimized with the B3LYP/6-31 g (d) basis set.

Figure 3. Experimental vs calculated shifts for 3$^{2-}$ from reduction with potassium.

Lithium or potassium counterions. The experimental and calculated shifts with the greatest agreement are given in Table 2. Figure 3 shows that there is good agreement between the experimental and calculated shift for the dianion reduced with potassium when the shift is calculated with the counterion and with solvent. The agreement is not quite as good for the dianion from reduction with lithium and is best when both solvent and counterion are included in the calculation; see the Supporting Information.

It is clear that there is an effect of the counterion and that the effect is strongest in the 1 position. This would suggest incomplete delocalization that left the greatest electron density on carbon-9. This would be consistent with a pattern of delocalization that preserved the stability of the benzene rings.

** Attempted Preparation of the Dianion of Dixaanthylene by Deprotonation. ** In attempts to prepare antiaromatic dianions by reduction in other systems, such as 5, we have found that there is a strong tendency for over-reduction to trianion radicals and tetraniions. Thus, we set out to determine a protocol for the formation of antiaromatic dianions through deprotonation. Deprotonation of 6, Scheme 2, would give an antiaromatic dianion whose $^1$H NMR spectrum is known, giving us proof of concept.

While deprotonation to give antiaromatic anions has been problematic, Wan et al. discovered that when deprotonation was done with irradiation, the photoexcited starting material was substantially more acidic. For example, the antiaromatic suberenyl anion could be formed with much greater ease, including with the use of very weak bases like D$_2$O. The ease of formation under irradiation is consistent with the observation that the excited states of antiaromatic species are aromatic.

Because we were interested in the direct observation of the dianion via $^1$H NMR spectroscopy, we chose to examine the deprotonation of 6 with an excess of strong base, $n$-butyllithium/TMEDA. We anticipated that formation of the dianion would be fairly smooth because the anion of xanthene was successfully generated via excited state deprotonation. Rather than getting deprotonation at the benzylic positions to give 3$^-$, the major product visible in reaction mixture from the quench of the mixture with CH$_3$I was 7, presumably from 6$^-$, see Scheme 2, although there was evidence of other methylated species in much smaller amounts as well as unreacted starting material; see the Supporting Information. The low isolated yield of 6 is unoptimized and reflects our primary interest in pure product. There is precedent for deprotonation of 9,9-dimethylxanthene to give the dianion, with deprotonation occurring ortho to the oxygen atoms.

In this case, there was no advantage to deprotonation under irradiation because the spectra of the irradiated and dark reaction mixtures were virtually identical; see the Supporting Information. The low isolated yield of 6 was unoptimized and reflects our primary interest in pure product. There is precedent for deprotonation of 9,9-dimethylxanthene to give the dianion, with deprotonation occurring ortho to the oxygen atoms.

We believe that the failure to observe deprotonation to the antiaromatic dianion supports an observation by Budac et al. in which they attributed the failure to deprotonate 8 to the geometry of the starting material. Deprotonation of 8, R = H, under photolysis resulted in significant deuteration exchange. However, the derivatives of 8 with R = CH$_3$ or Ph gave no deuteration exchange. When the proton in the 5-position of 8 was in the pseudoequatorial position, deprotonation would not result in "significant overlap of the developing $\pi$ orbital with the $\pi$ orbitals of the two benzene rings." This is in agreement with the known crystal structure of 6, in which one xanthyl ring occupies the axial position of the second xanthyl ring system, placing the C−H bond to be deprotonated in the pseudoequatorial position; see the Supporting Information for the crystal structure previously reported. While the similarity of geometry of 6 and 8 supports the failure of excited state deprotonation in both, the folded geometry of

Scheme 2. Tetra-deprotonation of 6

![Scheme 2. Tetra-deprotonation of 6](image)
8 and related species\textsuperscript{39} could also contribute to its reluctance to deprotonate.

\begin{center}
\includegraphics{figure}
\end{center}

In summary, we report the reduction of dixanthylidene 3 to an antiaromatic heterocyclic dianion, which was characterized through \textsuperscript{1}H NMR spectroscopy. Attempts to form the dianion by taking advantage of the potential excited state acidity of 6 were unsuccessful because the geometry of the starting material failed to allow adequate stabilization of the developing carbanion through its overlap with the \pi-system of the benzene rings.

\section*{EXPERIMENTAL SECTION}

\textbf{Reduction of 3 to 3\textsuperscript{2-}.} Dixanthylidene 3 was synthesized by the method of Ault.\textsuperscript{40} Neutral 3 was reduced with lithium following a variation of the procedure of Rabinowitz.\textsuperscript{41} A piece of fresh lithium wire was placed in the upper part of an extended NMR tube which contained ca. 10 mg of 3. THF-d\textsubscript{4} was transferred to the evacuated tube. The solution was degassed using the freeze-pump-thaw technique and flame-sealed. The solution was brought into contact with the lithium by inverting the tube. The inverted tube was then sonicated at 0°C to remove the oxidized layer on the lithium. Sonication continued until the solution turned to a deep red, ca. 8 h.

\textbf{Deprotonation of 6 and Formation of 7.} \textit{N,N,N',N'-Tetramethylene-}
dimethylethylenediamine (3.0 mL, 0.022 mol) was added to a solution of \textit{n-butyl lithium in cyclohexane} (2.2 M, 10 mL, 0.022 mol) giving a light yellow solution with a white gelatinous precipitate. This solution was then added to 6\textsuperscript{42} to give a brown solid. Purification through a silica plug followed by recrystallization from hexanes gave a white solid (0.012 g, 5.4% yield). Isolation procedures included irradiation of a yellow solution with a white gelatinous precipitate. This solution was added to 6\textsuperscript{42} with 0.16 mol methyl iodide at 0°C. The solution slowly turned deep red. 

\textbf{Computational Methods.} Geometries were optimized at B3LYP/6-31G(d) density functional theory levels with the Gaussian 98 and 03 program packages (see the Supporting Information). The chemical shifts were calculated at B3LYP/6-311+g(dp) using the GIAO approach with the Gaussian 98 or 03 program packages on the optimized geometries. The nucleus-independent chemical shifts (NICS(1))\textsuperscript{12,23} were obtained from the chemical shift tensor perpendicular to the ring for a dummy atom placed 1 Å above the center of each ring.

\section*{ASSOCIATED CONTENT}

\textbf{Supporting Information.} Geometries of \textit{1\textsuperscript{2-} and \textit{3\textsuperscript{2-} and related monoanions, details of the magnetic susceptibility calculations. complete \textsuperscript{1}H NMR spectral data for reduction of 3 with lithium and potassium, COSY spectrum from reduction with potassium, plots of experimental vs calculated shifts and tables of calculated shifts for dianions with and without counterion, with and without solvent, complete NMR characterization of 7 (\textsuperscript{1}H, \textsuperscript{13}C, COSY, HMQC, HMBC data), and \textsuperscript{1}H NMR spectra of irradiated and nonirradiated deprotonation. \textsuperscript{1}H NMR spectrum of the reaction mixture from quench with methyl iodide, reported \textsuperscript{1}H NMR data for 4, literature data on the crystal structure of 6, Cartesian coordinates and total energies for \textit{1\textsuperscript{2-}, \textit{2\textsuperscript{2-}, \textit{K2, Li2, and related monoanions, complete citations for Gaussian 03 and 09. This material is available free of charge via the Internet at http://pubs.acs.org.}

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\section*{ACKNOWLEDGMENT}

We thank the Welch Foundation (Grant No. W-794) and the National Science Foundation (Grant Nos. CHE-0242227 and CHE-0948445) for support of this work.

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